



<u>script</u>	<u>board</u>
<p><u>Introduction</u></p> <ul style="list-style-type: none"> - myself - paula --- No TA in CE3305-2014-2.2 - web server write url - why <u>not</u> ttu server - distribute syllabus Electronic distribute only - show required textbook - state: <u>really will need to read the book!</u> Use the e-book if possible, I did bring a physical book - explain <u>why</u> roll sheet - distribute roll sheet - questions? 	<p>T.G. Cleveland</p> <p>P.A. --- Monaco ---</p> <p style="text-align: right;"><i>need this character!</i></p> <p>www.rtfmps.com/ce3305/ give the userid and password to class to get by the paywall</p> <p><u>Syllabus</u></p> <ul style="list-style-type: none"> - objectives - schedule - grading <ul style="list-style-type: none"> - quizzes - exercises - exams - ABET PEO <p>Roll sheet</p>
<ul style="list-style-type: none"> - Exercise format - Solution Process - Explain purpose of format & process <p>Summer class is really accelerated Need to do the exercises to be able to complete the quizzes. Quizzes are begin each meeting, 10 minutes -- no adjustments. They are part of the course letter grade, and are essentially an on-going exam. Two exams (if this were a normal semester would be in class) are take home. One exam in-class.</p>	<p><u>Exercises</u></p> <ul style="list-style-type: none"> - Practice at problem solving - Format - Process <p><u>Purpose:</u> (1) Develop a systematic method to solve closed engineering/science problems.</p> <p>(2) Develop "muscle memory" to apply ^{the} process</p> <p>(3) Become accustomed to documenting effort.</p>



Script

Process

→ Essentially the scientific method with hypothesis replaced by governing equations.

→ Usually works, worst case identifies things that need clarification

I have this out of order:

- 1) state problem
- 2) sketch, knowns
- 3) unknowns, find
- 4) solution -- identify governing equations in advance and as used.
- 5) validate/discuss results

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state the problem, sketch

Identify "knowns"

Identify governing equations, assumptions, and principles

List "unknowns"

Solve for "unknowns"

Validate/discuss results

- sensitivity
- effect of relaxing an assumption

Format

→ How we (i) want exercises to appear.

(1) Each problem starts on a new page!

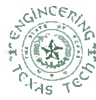
Latex / clearpage

(2) Upper right hand every page!

Name Date
Sheet X of Y

(3) Use engineering grid paper for "final" submission

	NAME	DATE
		2 OF X
1)	PROBLEM STATEMENT	
2)	SKETCH OF SITUATION / SYSTEM (FBD IF APPROPRIATE)	
3)	KNOWN	
	⋮	
4)	GOVERNING EQUATION(S)	
	⋮	
5)	UNKNOWN(S) / FIND....	
6)	SOLUTION	
7)	DISCUSSION	



Script

Example to illustrate format

Remark as example proceeds:

OK to make mistake -

Use ~~strike through~~ and continue.

OK to use guide arrows

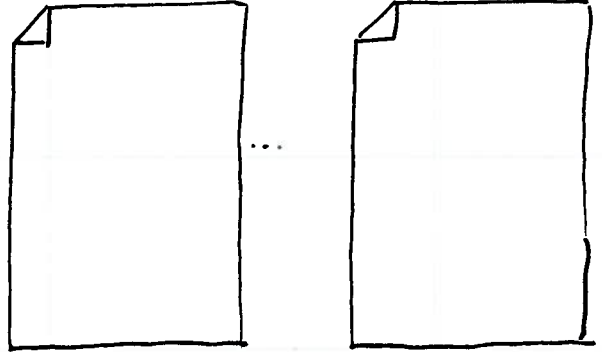


Remind to study notes

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Follow EXAMPLE ~~HERE SHEET~~ ATTACHED AT END.

DRAW AS:



INSERT EXAMPLE HERE

Quiz 1 is essentially this problem.

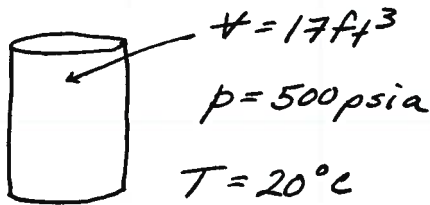
PROBLEM STATEMENT:

FIND THE TOTAL WEIGHT OF A 17ft^3 TANK OF NITROGEN AT 500 psia.

THE VESSEL ITSELF WEIGHS 50 lbf.

THE TEMPERATURE IS 20°C

SKETCH (FBD EQUIVALENT):



KNOWN

$$V = 17\text{ft}^3$$

$$p = 500\text{psia}$$

$$T = 20^\circ\text{C}$$

$$W_{\text{TANK}} = 50\text{lbf}$$

GOVERNING EQUATIONS:

IDEAL GAS LAW $pV = \frac{m}{M}RT$ (ASSUME IDEAL GAS)
LAW APPLIES

$$W_{\text{TOTAL}} = W_{\text{TANK}} + W_{\text{GAS}}$$

UNKNOWN: (FIND)

$$W_{\text{GAS}}$$

SOLUTION

$$W_{gas} = m_{gas} \cdot g$$

WHERE R IS UNIVERSAL GAS CONSTANT

$$m_{gas} = \frac{pV M}{RT}$$

M IS MOLECULAR WEIGHT OF NITROGEN

$$R = 0.0821 \frac{L \cdot atm}{K \cdot mol}$$

$$M_N = 14.007 g/mol \quad \therefore \quad M_{N_2} = 28.014 g/mol$$

GAS IS N_2 ; EVERY MOLE IS 28g!

NEED V IN L

NEED T IN $^{\circ}K$

NEED p IN ATM

$$V = 17 ft^3 \cdot \frac{1m}{3.28ft} \cdot \frac{1m}{3.28ft} \cdot \frac{1m}{3.28ft} \cdot \frac{1000L}{1m^3} = \underline{481.76 L}$$

← Volume

$$^{\circ}C + 273 = ^{\circ}K$$

$$T = 20^{\circ}C + 273 = \underline{293^{\circ}K}$$

← TEMP.

$$p = 500 psi \cdot \frac{1 atm}{14.7 psi} = \underline{34.01 atm}$$

← PRESSURE

NOW SOLVE FOR MASS OF GAS

$$* m_{gas} = \frac{(34.01 atm)(481.76 L)(28.014 g/mol)}{(0.0821 \frac{L \cdot atm}{K \cdot mol})(293 K)} = \underline{19071.5 g N_2}$$

← MASS GAS

NOW APPLY W_{gas} TO FIND WEIGHT OF GAS

$$W_{gas} = 19071.5g \frac{kg}{1000g} \cdot 9.81m/s^2 = \underline{187.09 N} \leftarrow \text{WEIGHT GAS}$$

NOW CONVERT VESSEL WEIGHT TO NEWTONS

$$W_{TANK} = 50 lbf \cdot \frac{1kg}{2.2lbf} \cdot 9.8m/s^2 = \underline{222.72 N} \leftarrow \text{WEIGHT TANK}$$

TOTAL WEIGHT

$$W_T = W_{gas} + W_{tank}$$

$$= 187.09N + 222.72N = \underline{409.81N} \leftarrow \text{TOTAL WEIGHT IN NEWTONS}$$

PROBABLY MAKES SENSE TO CONVERT TO LBS

$$409.8N = 409.8 \frac{kg \cdot m}{s^2} \cdot \frac{1}{9.8m/s^2} \cdot \frac{2.2lbf}{1kg} = \underline{91.99 lbf} \leftarrow \text{TOTAL WEIGHT IN LBS.}$$

DISCUSSION

SEVERAL INTERMEDIATE VALUES ARE REQUIRED, ESP. R AND M.

NEED TO RECOGNIZE THAT N_2 HAS TWO ATOMS PER MOLE, THUS M_{N_2} IS TWICE M_N

NEED TABLE OF ELEMENTS TO LOOK UP MOLECULAR WEIGHT NITROGEN (ATTACHED AS EXHIBIT 1)

FINAL ANSWER TOO MANY SIGNIFICANT FIGURES;
409N AND 92lbf WOULD BE MORE MEANINGFUL.

REMEMBER TO REPORT UNITS WITH ANSWER!

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EXAMPLE FORMAT

Periodic Table of the Elements



atomic number → 14
atomic weight → 28.09
name → Silicon

symbol: **Si**

- black solid
- blue liquid
- red gas
- synthetically prepared
- metal

Legend:

- alkali metals
- alkaline earth metals
- transitional metals
- other metals
- nonmetals
- noble gases

1	H Hydrogen 1.008	2	He Helium 4.003
3	Li Lithium 6.941	4	Be Beryllium 9.012
11	Na Sodium 22.990	12	Mg Magnesium 24.305
19	K Potassium 39.102	20	Ca Calcium 40.078
37	Rb Rubidium 85.468	38	Sr Strontium 87.62
55	Cs Cesium 132.905	56	Ba Barium 137.327
87	Fr Francium (223)	88	Ra Radium (226)
21	Sc Scandium 44.956	22	Ti Titanium 47.88
39	Y Yttrium 88.906	40	Zr Zirconium 91.224
57	La Lanthanum 138.905	72	Hf Hafnium 178.49
89	Ac Actinium 227.03	104	Rf Rutherfordium (261)
23	V Vanadium 50.942	24	Cr Chromium 51.996
41	Nb Niobium 92.906	42	Mo Molybdenum 95.94
73	Ta Tantalum 180.948	74	W Tungsten 183.85
105	Hf Hafnium (262)	106	Sg Seaborgium (266)
25	Mn Manganese 54.938	26	Fe Iron 55.845
43	Tc Technetium (98)	44	Ru Ruthenium 101.07
75	Re Rhenium 186.21	76	Os Osmium 190.20
107	Bh Bohrium (262)	108	Hs Hassium (265)
27	Co Cobalt 58.933	28	Ni Nickel 58.708
45	Rh Rhodium 102.91	46	Pd Palladium 106.40
77	Ir Iridium 192.22	78	Pt Platinum 195.08
109	Mt Meitnerium (268)	110	Ds Darmstadtium (271)
29	Cu Copper 63.546	30	Zn Zinc 65.37
47	Ag Silver 107.868	48	Cd Cadmium 112.411
79	Au Gold 196.967	80	Hg Mercury 200.59
111	Rg Roentgenium (272)	112	Cn Copernicium (277)
5	B Boron 10.811	6	C Carbon 12.011
13	Al Aluminum 26.982	14	Si Silicon 28.086
31	Ga Gallium 69.723	32	Ge Germanium 72.64
49	In Indium 114.818	50	Sn Tin 118.710
81	Tl Thallium 204.38	82	Pb Lead 207.19
113	Nh Nihonium (284)	114	Fl Flerovium (289)
7	N Nitrogen 14.007	8	O Oxygen 15.999
15	P Phosphorus 30.974	16	S Sulfur 32.06
33	As Arsenic 74.922	34	Se Selenium 78.96
51	Sb Antimony 121.757	52	Te Tellurium 127.60
83	Bi Bismuth 208.98	84	Po Polonium (209)
115	Mc Moscovium (288)	116	Lv Livermorium (293)
9	F Fluorine 18.998	10	Ne Neon 20.18
17	Cl Chlorine 35.453	18	Ar Argon 39.95
35	Br Bromine 79.904	36	Kr Krypton 83.80
53	I Iodine 126.905	54	Xe Xenon 131.30
85	At Astatine (210)	86	Rn Radon (222)
117	Ts Tennessine (294)	118	Og Oganesson (294)
58	Ce Cerium 140.12	59	Pr Praseodymium 140.91
90	Th Thorium 232.04	91	Pa Protactinium 231.04
60	Nd Neodymium 144.24	61	Pm Promethium (145)
92	U Uranium 238.03	93	Np Neptunium 237.05
62	Sm Samarium 150.4	63	Eu Europium 151.96
94	Pu Plutonium (244)	95	Am Americium (243)
64	Gd Gadolinium 157.25	65	Tb Terbium 158.93
96	Cm Curium (247)	97	Bk Berkelium (247)
66	Dy Dysprosium 162.50	67	Ho Holmium 164.93
98	Cf Californium (251)	99	Es Einsteinium (252)
68	Er Erbium 167.26	69	Tm Thulium 168.93
100	Fm Fermium (257)	101	Md Mendelevium (260)
70	Yb Ytterbium 173.04	71	Lu Lutetium 174.97
102	No Nobelium (262)	103	Lr Lawrencium (262)

http://oxford-labs.com/wp-content/uploads/2009/04/periodic-table.jpg

url

DO THE EXAMPLE BEFORE START THIS SECTION

What is fluid mechanics?

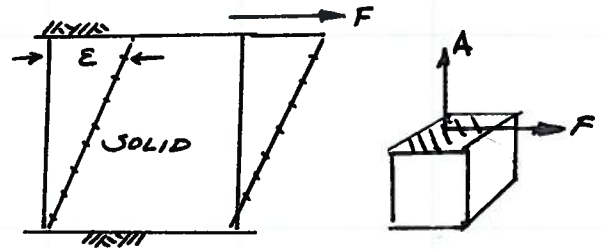
A solid deforms, but not continuously.

When stress is applied, deformation is proportional to stress

Deformation (strain) is some function of stress (τ).

Assuming solid does not fail, then deformation is largely time invariant

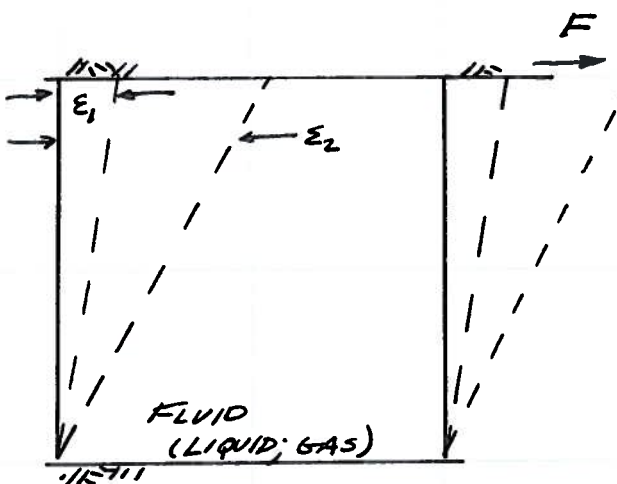
(ϵ does not change with time)



$$\epsilon \propto \frac{F}{A} = \tau \quad \text{"SHEAR STRESS"}$$

$$\epsilon = f(\tau) \quad \frac{d\epsilon}{dt} = 0$$

Time ↗

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<p>A fluid (liquid or gas) deforms <u>continuously</u>; as long as stress is applied, the fluid will continue to deform.</p> <p>Deformation (strain) is proportional to stress.</p> <p>Strain is <u>time</u> variant.</p> <p>(look ahead pg 36 - note) similarity of sketches</p>	 $\epsilon \propto \frac{F}{A} = \gamma$ $\epsilon \propto t \Rightarrow \epsilon = g(\gamma, t)$ $\frac{d\epsilon}{dt} \neq 0$
<p>What is fluid mechanics?</p> <p>Compare to pg 3 definition</p> <p>What kind of things are fluids?</p> <ul style="list-style-type: none"> • water (at & P = ??) • air • gasoline • paraffin at (60°F, 190°F)? • window glass? • ice cream? • paint? 	<p>Fluid mechanics is the study of motion, deformation, momentum, energy, and related properties of materials that cannot resist shear stress</p> <p>Liquids or gasses.</p>

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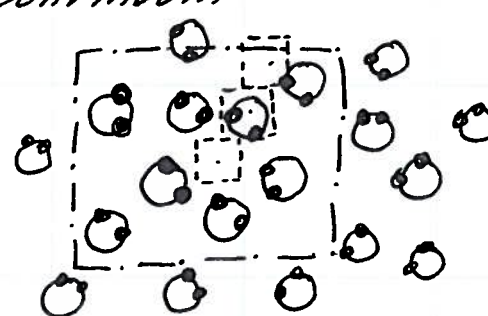
Calculus of fluid mechanics uses continuously differentiable functions (usually) and implicitly assumes a fluid is a continuum

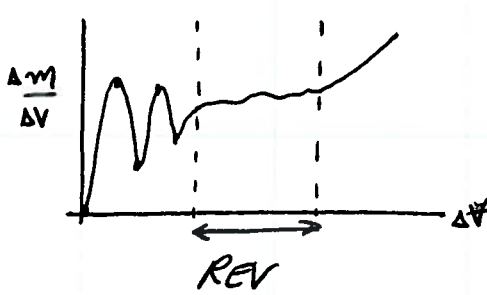
Usual definition is that a mass density can be defined - and is usually assumed to be controlled by the sampling volume ΔV .

The lowest usable value ΔV is called a representative elementary volume (REV)

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Continuum



$$\rho = \frac{\Delta m}{\Delta V}$$


Concept of fluid parcel is useful especially when studying motion in a Lagrangian coordinate system.

A fluid parcel is a quantity of fluid with fixed mass identity (usually mass)

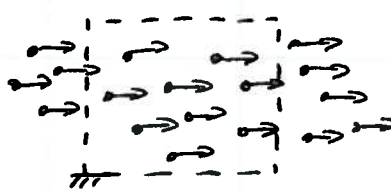
The volume of parcel can remain finite³ or be infinitesimal.

When infinitesimal, they are called fluid particles


2 Fluid parcels used in air quality models where parcels represent cubic kilometers.

3 Some models represent parcels as "drops" (i.e. $1 \cdot 10^6$ drops = 1 watershed inch) - these are particles

Fluid parcel (particle)



Eulerian



Lagrangian

Fixed identity (usually mass)

